ARKANSAS CHEMISTRY · UNIT 3 · 10 QUESTIONS

Stoichiometry and Solutions — Practice Test

This practice test was made by AI and there may be mistakes. If a question or answer looks wrong, use the Flag a mistake button to let Brett know so he can fix it.

1 **A reaction makes 2 moles of product from 4 moles of reactant A in a 1 to 1 ratio, if only 1 mole of A is available, how many moles of product form**

ANSWER _____

Understand It with Pulse | Flag a mistake

2 **How many particles are in 2 moles, in multiples of Avogadro number, give the number only**

ANSWER _____

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3 **A hydrate contains water molecules within its crystal structure, and heating it usually**

- A Adds water
- B Changes its elements
- C Makes it ionic only
- D Drives off the water

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4 **Colligative properties depend on the number of solute particles, not their identity, for example boiling point**

- A Elevation
- B Depression only
- C Colour
- D Density only

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5 **Adding salt to water raises its boiling point and lowers its**

A Density always

B Freezing point

C Volume always

D Mass

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6 **An electrolyte solution conducts electricity because it contains**

A Free electrons only

B No particles

C Free ions

D Only molecules

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7 **A nonelectrolyte solution does not conduct electricity because its solute forms**

A Ions that are fixed

B Metals

C Gases

D Molecules, not ions, in solution

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8 **Molarity changes with temperature because volume changes, while molality uses mass of solvent instead. How does molality behave as temperature changes?**

A Independent of temperature changes in volume

B Dependent on volume

C Equal to molarity always

D Measured in litres

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9 **Dilution equation $M_1 V_1$ equals $M_2 V_2$ can be used when moles of solute stay**

A Increasing

B Constant during dilution

C Decreasing

D Zero

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10 **A precipitate in a reaction can be separated from solution by**

A Distillation only

B Melting

C Filtration

D Evaporation only

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ANSWER KEY

1.	1	**5.**	B	**9.**	B
2.	2	**6.**	C	**10.**	C
3.	D	**7.**	D		
4.	A	**8.**	A		

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