ARKANSAS CHEMISTRY · UNIT 4 · 10 QUESTIONS

Gases, Thermodynamics, and Equilibrium — Practice Test

This practice test was made by AI and there may be mistakes. If a question or answer looks wrong, use the Flag a mistake button to let Brett know so he can fix it.

1 **Convert 27 degrees Celsius to kelvin**

ANSWER

Understand It with Pulse | Flag a mistake

2 **A mixture has partial pressures 2 atmospheres oxygen and 3 atmospheres nitrogen, what is the total pressure in atmospheres**

ANSWER

Understand It with Pulse | Flag a mistake

3 **A solid absorbs heat as it melts, while its temperature stays the same. The energy absorbed during this change of state is called the**

- A heat of fusion
- B specific heat
- C heat of vaporisation
- D thermal expansion

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4 **Specific heat capacity is energy needed to raise temperature of one gram by**

- A One degree Celsius
- B One hundred degrees
- C One kelvin only at high temperature
- D One mole only

Understand It with Pulse | Flag a mistake

5 **A calorimeter measures heat change of a reaction by measuring temperature change of**

- A The gas only
- B Surrounding water or solution
- C The catalyst
- D The container only

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6 **Hess law states total enthalpy change is independent of the**

- A Temperature
- B Pressure
- C Path taken
- D Catalyst

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7 **Spontaneous does not mean fast, it means a process can occur without**

- A Any energy change
- B A catalyst
- C Equilibrium
- D Continuous outside energy input

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8 **Vapour pressure increases with temperature because more molecules have enough energy to**

- A Escape the liquid surface
- B Freeze
- C Dissolve
- D Condense

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9 **Boiling occurs when vapour pressure equals**

A Zero

B Atmospheric pressure

C Partial pressure

D Water pressure only

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10 **Real gases deviate most from ideal behaviour at high pressure and low temperature because particle volume and attractions become**

A Negligible

B Zero

C Significant

D Constant

Understand It with Pulse | Flag a mistake

ANSWER KEY

1.	300	**5.**	B	**9.**	B
2.	5	**6.**	C	**10.**	C
3.	A	**7.**	D		
4.	A	**8.**	A		

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