ALBERTA CHEMISTRY 20 · UNIT C · 5 QUESTIONS

Matter as Solutions, Acids and Bases — Practice Test

This practice test was made by AI and there may be mistakes. If a question or answer looks wrong, use the Flag a mistake button to let Brett know so he can fix it.

1 **A student dissolves 1.5 mol of solute in enough water to make 500 mL of solution. The concentration of the resulting solution is**

- A 0.33 mol/L
- B 0.75 mol/L
- C 3.0 mol/L
- D 1.5 mol/L

Understand It with Pulse | Flag a mistake

2 **A small crystal of solute is added to a solution, and the crystal dissolves completely. Before the crystal was added, the solution was**



A unsaturated

B saturated

C supersaturated

D concentrated

Understand It with Pulse | Flag a mistake

3 **Which of the following hydronium ion concentrations represents the most acidic solution?**

A 1.0×10^{-2} mol/L

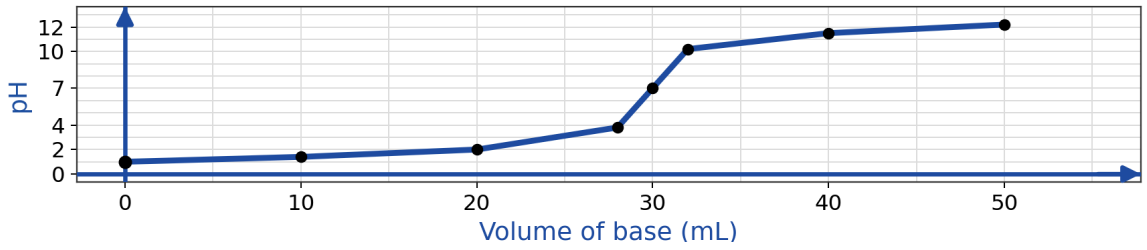
B 1.0×10^{-5} mol/L

C 1.0×10^{-7} mol/L

D 1.0×10^{-11} mol/L

Understand It with Pulse | Flag a mistake

4 **The titration curve below shows the pH as sodium hydroxide solution is added to an acid. At what volume of base added is the equivalence point reached?**



- **A** 20 mL
- **B** 30 mL
- **C** 40 mL
- **D** 50 mL

Understand It with Pulse | Flag a mistake

5 **A 200 mL sample of 0.50 mol/L sodium chloride solution is diluted with water to a final volume of 1.0 L. The concentration of the diluted solution is**

- A 2.5 mol/L
- B 0.25 mol/L
- C 0.10 mol/L
- D 0.50 mol/L

Understand It with Pulse | Flag a mistake

ANSWER KEY

1. C
2. A
3. A
4. B
5. C

© 2026 Brett Robbins. All Rights Reserved.