CONNECTICUT CHEMISTRY · UNIT 3 · 10 QUESTIONS

Stoichiometry and Solutions — Practice Test

This practice test was made by AI and there may be mistakes. If a question or answer looks wrong, use the Flag a mistake button to let Brett know so he can fix it.

1 **Which statement about Solution is correct**

- A the substance dissolved in a solution
- B the substance doing the dissolving
- C a uniform mixture of solute and solvent
- D the amount of solute per amount of solution

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2 **Which statement about Solute is correct**

- A the substance doing the dissolving
- B the amount of solute per amount of solution
- C moles of solute per liter of solution
- D the substance dissolved in a solution

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3 **Which statement about Solvent is correct**

A the substance doing the dissolving

B the amount of solute per amount of solution

C moles of solute per liter of solution

D adding solvent to lower concentration

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4 **Which statement about Concentration is correct**

A moles of solute per liter of solution

B the amount of solute per amount of solution

C adding solvent to lower concentration

D a substance producing hydrogen ions in water

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5 **Which statement about Molarity is correct**

A adding solvent to lower concentration

B a substance producing hydrogen ions in water

C moles of solute per liter of solution

D a substance producing hydroxide ions in water

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6 **A solution has 2 moles of solute in 1 liter. Its molarity is**

ANSWER _____

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7 **A solution has 6 moles of solute in 2 liters. Its molarity is**

ANSWER _____

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8 **A reaction has a theoretical yield of 100 grams and an actual yield of 80 grams. Its percent yield is**

ANSWER _____

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9 **A balanced equation uses 2 moles of hydrogen per 1 mole of oxygen. For 4 moles of hydrogen, the moles of oxygen needed are**

ANSWER _____

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10 **A sample has molar mass 18 grams per mole. The mass of 3 moles is**

ANSWER _____

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ANSWER KEY

1.	C	**5.**	C	**9.**	2
2.	D	**6.**	2	**10.**	54
3.	A	**7.**	3		
4.	B	**8.**	80		

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