DELAWARE CHEMISTRY · UNIT 2 · 10 QUESTIONS

Bonding and Chemical Reactions — Practice Test

This practice test was made by AI and there may be mistakes. If a question or answer looks wrong, use the Flag a mistake button to let Brett know so he can fix it.

1 **Which term is defined as follows: a reaction where two compounds exchange partners?**

A ionic bond

B covalent bond

C double replacement

D polar bond

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2 **Which of the following is an example of double replacement?**

- A sodium chloride held by opposite charges
- B the bonds inside a water molecule
- C the oxygen hydrogen bond in water
- D two salt solutions forming a precipitate

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3 **Which statement best describes double replacement?**

- A double replacement often produces a solid, gas, or water
- B ionic compounds form crystal lattices
- C covalent bonds hold atoms together in molecules
- D polar bonds arise from differences in electronegativity

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4 **Which term is defined as follows: an insoluble solid formed in a solution?**

A ionic bond

B covalent bond

C polar bond

D precipitate

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5 **Which of the following is an example of precipitate?**

A a cloudy solid appearing in a beaker

B sodium chloride held by opposite charges

C the bonds inside a water molecule

D the oxygen hydrogen bond in water

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6 **Which statement best describes precipitate?**

A ionic compounds form crystal lattices

B a precipitate signals that a reaction occurred

C covalent bonds hold atoms together in molecules

D polar bonds arise from differences in electronegativity

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7 **Which term is defined as follows: a substance that speeds a reaction without being used up?**

A catalyst

B ionic bond

C covalent bond

D polar bond

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8 **Which of the following is an example of catalyst?**

A sodium chloride held by opposite charges

B an enzyme in a cell

C the bonds inside a water molecule

D the oxygen hydrogen bond in water

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9 **Which statement best describes catalyst?**

A ionic compounds form crystal lattices

B covalent bonds hold atoms together in molecules

C a catalyst lowers activation energy

D polar bonds arise from differences in electronegativity

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10 **Which term is defined as follows: the minimum energy needed to start a reaction?**

A ionic bond

B activation energy

C covalent bond

D polar bond

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ANSWER KEY

1.	C	**5.**	A	**9.**	C
2.	D	**6.**	B	**10.**	B
3.	A	**7.**	A		
4.	D	**8.**	B		

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