DELAWARE CHEMISTRY · UNIT 3 · 25 QUESTIONS

Stoichiometry and Solutions — Practice Test

This practice test was made by AI and there may be mistakes. If a question or answer looks wrong, use the Flag a mistake button to let Brett know so he can fix it.

1 **Which term is defined as follows: the amount containing Avogadro's number of particles?**

- A mole
- B Avogadro's number
- C molar mass
- D stoichiometry

Understand It with Pulse | Flag a mistake

2 **Which of the following is an example of mole?**

- A the number of atoms in 12 grams of carbon 12
- B a dozen eggs as a counting unit
- C 18 grams per mole for water
- D finding grams of product from grams of reactant

Understand It with Pulse | Flag a mistake

3 **Which statement best describes mole?**

A Avogadro's number links moles to particle counts

B molar mass is found from atomic masses

C a mole counts particles, not mass directly

D stoichiometry uses mole ratios from balanced equations

Understand It with Pulse | Flag a mistake

4 **Which term is defined as follows: about 6.02 times 10 to the 23rd particles per mole?**

A mole

B Avogadro's number

C molar mass

D stoichiometry

Understand It with Pulse | Flag a mistake

5 **Which of the following is an example of Avogadro's number?**

- A a dozen eggs as a counting unit
- B 18 grams per mole for water
- C the number of atoms in 12 grams of carbon 12
- D finding grams of product from grams of reactant

Understand It with Pulse | Flag a mistake

6 **Which statement best describes Avogadro's number?**

- A a mole counts particles, not mass directly
- B molar mass is found from atomic masses
- C stoichiometry uses mole ratios from balanced equations
- D Avogadro's number links moles to particle counts

Understand It with Pulse | Flag a mistake

7 **Which term is defined as follows: the mass of one mole of a substance?**

A mole

B Avogadro's number

C molar mass

D stoichiometry

Understand It with Pulse | Flag a mistake

8 **Which of the following is an example of molar mass?**

A a dozen eggs as a counting unit

B the number of atoms in 12 grams of carbon 12

C finding grams of product from grams of reactant

D 18 grams per mole for water

Understand It with Pulse | Flag a mistake

9 **Which statement best describes molar mass?**

A molar mass is found from atomic masses

B a mole counts particles, not mass directly

C Avogadro's number links moles to particle counts

D stoichiometry uses mole ratios from balanced equations

Understand It with Pulse | Flag a mistake

10 **Which term is defined as follows: the calculation of quantities in chemical reactions?**

A mole

B Avogadro's number

C molar mass

D stoichiometry

Understand It with Pulse | Flag a mistake

11 **Which of the following is an example of stoichiometry?**

- A finding grams of product from grams of reactant
- B a dozen eggs as a counting unit
- C the number of atoms in 12 grams of carbon 12
- D 18 grams per mole for water

Understand It with Pulse | Flag a mistake

12 **Which statement best describes stoichiometry?**

- A a mole counts particles, not mass directly
- B stoichiometry uses mole ratios from balanced equations
- C Avogadro's number links moles to particle counts
- D molar mass is found from atomic masses

Understand It with Pulse | Flag a mistake

13 **Which term is defined as follows: the reactant that runs out first and limits product?**

A limiting reactant

B mole

C Avogadro's number

D molar mass

Understand It with Pulse | Flag a mistake

14 **Which of the following is an example of limiting reactant?**

A a dozen eggs as a counting unit

B the ingredient that runs out in a recipe

C the number of atoms in 12 grams of carbon 12

D 18 grams per mole for water

Understand It with Pulse | Flag a mistake

15 **Which statement best describes limiting reactant?**

- A a mole counts particles, not mass directly
- B Avogadro's number links moles to particle counts
- C the limiting reactant determines how much product can form
- D molar mass is found from atomic masses

Understand It with Pulse | Flag a mistake

16 **Which term is defined as follows: a reactant left over after a reaction?**

- A mole
- B excess reactant
- C Avogadro's number
- D molar mass

Understand It with Pulse | Flag a mistake

17 **Which of the following is an example of excess reactant?**

A a dozen eggs as a counting unit

B the number of atoms in 12 grams of carbon 12

C extra flour after baking

D 18 grams per mole for water

Understand It with Pulse | Flag a mistake

18 **Which statement best describes excess reactant?**

A a mole counts particles, not mass directly

B Avogadro's number links moles to particle counts

C molar mass is found from atomic masses

D an excess reactant is not fully used up

Understand It with Pulse | Flag a mistake

19 **Which term is defined as follows: actual yield divided by theoretical yield, times 100?**

A mole

B Avogadro's number

C percent yield

D molar mass

Understand It with Pulse | Flag a mistake

20 **Which of the following is an example of percent yield?**

A a dozen eggs as a counting unit

B the number of atoms in 12 grams of carbon 12

C 18 grams per mole for water

D 80 percent when 80 grams form of a possible 100

Understand It with Pulse | Flag a mistake

21 **Which statement best describes percent yield?**

A percent yield compares what formed with what was possible

B a mole counts particles, not mass directly

C Avogadro's number links moles to particle counts

D molar mass is found from atomic masses

Understand It with Pulse | Flag a mistake

22 **Which term is defined as follows: a uniform mixture with a solute dissolved in a solvent?**

A mole

B Avogadro's number

C molar mass

D solution

Understand It with Pulse | Flag a mistake

23 **Which of the following is an example of solution?**

- A salt water
- B a dozen eggs as a counting unit
- C the number of atoms in 12 grams of carbon 12
- D 18 grams per mole for water

Understand It with Pulse | Flag a mistake

24 **Which statement best describes solution?**

- A a mole counts particles, not mass directly
- B a solution looks the same throughout
- C Avogadro's number links moles to particle counts
- D molar mass is found from atomic masses

Understand It with Pulse | Flag a mistake

25 **Which term is defined as follows: the substance dissolved in a solution?**

A solute

B mole

C Avogadro's number

D molar mass

Understand It with Pulse | Flag a mistake

ANSWER KEY

1. A
2. B
3. C
4. B
5. C
6. D
7. C
8. D
9. A
10. D
11. A
12. B
13. A
14. B
15. C
16. B
17. C
18. D
19. C
20. D
21. A
22. D
23. A
24. B
25. A

© 2026 Brett Robbins. All Rights Reserved.