DELAWARE CHEMISTRY · UNIT 4 · 25 QUESTIONS

Gases, Thermodynamics, and Equilibrium — Practice Test

This practice test was made by AI and there may be mistakes. If a question or answer looks wrong, use the Flag a mistake button to let Brett know so he can fix it.

1 **Which term is defined as follows: force applied per unit area?**

- A pressure
- B Boyle's law
- C Charles's law
- D ideal gas

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2 **Which of the following is an example of pressure?**

A squeezing a balloon raises its pressure

B a tire pressing on the road

C a balloon expanding when heated

D a model gas for calculations

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3 **Which statement best describes pressure?**

A Boyle's law is an inverse relationship

B Charles's law uses kelvin temperatures

C gas pressure comes from particle collisions with walls

D real gases behave most ideally at high temperature and low pressure

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4 **Which term is defined as follows: pressure times volume stays constant at fixed temperature?**

A pressure

B Boyle's law

C Charles's law

D ideal gas

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5 **Which of the following is an example of Boyle's law?**

A a tire pressing on the road

B a balloon expanding when heated

C squeezing a balloon raises its pressure

D a model gas for calculations

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6 **Which statement best describes Boyle's law?**

- A gas pressure comes from particle collisions with walls
- B Charles's law uses kelvin temperatures
- C real gases behave most ideally at high temperature and low pressure
- D Boyle's law is an inverse relationship

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7 **Which term is defined as follows: volume divided by temperature stays constant at fixed pressure?**

- A pressure
- B Boyle's law
- C Charles's law
- D ideal gas

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8 **Which of the following is an example of Charles's law?**

- A a tire pressing on the road
- B squeezing a balloon raises its pressure
- C a model gas for calculations
- D a balloon expanding when heated

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9 **Which statement best describes Charles's law?**

- A Charles's law uses kelvin temperatures
- B gas pressure comes from particle collisions with walls
- C Boyle's law is an inverse relationship
- D real gases behave most ideally at high temperature and low pressure

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10 **Which term is defined as follows: a gas whose particles have negligible volume and no attractions?**

A pressure

B Boyle's law

C Charles's law

D ideal gas

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11 **Which of the following is an example of ideal gas?**

A a model gas for calculations

B a tire pressing on the road

C squeezing a balloon raises its pressure

D a balloon expanding when heated

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12 **Which statement best describes ideal gas?**

A gas pressure comes from particle collisions with walls

B real gases behave most ideally at high temperature and low pressure

C Boyle's law is an inverse relationship

D Charles's law uses kelvin temperatures

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13 **Which term is defined as follows: the model explaining gas behavior from particle motion?**

A kinetic molecular theory

B pressure

C Boyle's law

D Charles's law

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14 **Which of the following is an example of kinetic molecular theory?**

- **A** a tire pressing on the road
- **B** gas particles moving randomly and colliding
- **C** squeezing a balloon raises its pressure
- **D** a balloon expanding when heated

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15 **Which statement best describes kinetic molecular theory?**

- **A** gas pressure comes from particle collisions with walls
- **B** Boyle's law is an inverse relationship
- **C** kinetic theory links temperature to average kinetic energy
- **D** Charles's law uses kelvin temperatures

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16 **Which term is defined as follows: the temperature at which particle motion is lowest?**

A pressure

B absolute zero

C Boyle's law

D Charles's law

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17 **Which of the following is an example of absolute zero?**

A a tire pressing on the road

B squeezing a balloon raises its pressure

C 0 kelvin

D a balloon expanding when heated

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18 **Which statement best describes absolute zero?**

A gas pressure comes from particle collisions with walls

B Boyle's law is an inverse relationship

C Charles's law uses kelvin temperatures

D absolute zero cannot be reached in practice

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19 **Which term is defined as follows: heat content of a system at constant pressure?**

A pressure

B Boyle's law

C enthalpy

D Charles's law

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20 **Which of the following is an example of enthalpy?**

- A a tire pressing on the road
- B squeezing a balloon raises its pressure
- C a balloon expanding when heated
- D heat absorbed in a reaction

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21 **Which statement best describes enthalpy?**

- A a negative enthalpy change means heat is released
- B gas pressure comes from particle collisions with walls
- C Boyle's law is an inverse relationship
- D Charles's law uses kelvin temperatures

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22 **Which term is defined as follows: a reaction that releases heat?**

A pressure

B Boyle's law

C Charles's law

D exothermic reaction

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23 **Which of the following is an example of exothermic reaction?**

A burning wood giving off warmth

B a tire pressing on the road

C squeezing a balloon raises its pressure

D a balloon expanding when heated

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24 **Which statement best describes exothermic reaction?**

A gas pressure comes from particle collisions with walls

B exothermic reactions warm their surroundings

C Boyle's law is an inverse relationship

D Charles's law uses kelvin temperatures

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25 **Which term is defined as follows: a reaction that absorbs heat?**

A endothermic reaction

B pressure

C Boyle's law

D Charles's law

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ANSWER KEY

1. A
2. B
3. C
4. B
5. C
6. D
7. C
8. D
9. A
10. D
11. A
12. B
13. A
14. B
15. C
16. B
17. C
18. D
19. C
20. D
21. A
22. D
23. A
24. B
25. A

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