DELAWARE CHEMISTRY · FINAL EXAM · 50 QUESTIONS

Final Exam — Practice Test

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1

Which term is defined as follows: the smallest unit of an element that keeps its properties?

A atom

B element

C compound

D mixture

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2 **Which term is defined as follows: a bond formed by transfer of electrons between a metal and a nonmetal?**

A ionic bond

B covalent bond

C polar bond

D hydrogen bond

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3 **Which of the following is an example of mole?**

A the number of atoms in 12 grams of carbon 12

B a dozen eggs as a counting unit

C 18 grams per mole for water

D finding grams of product from grams of reactant

Understand It with Pulse | Flag a mistake

4 **Which of the following is an example of pressure?**

- A squeezing a balloon raises its pressure
- B a tire pressing on the road
- C a balloon expanding when heated
- D a model gas for calculations

Understand It with Pulse | Flag a mistake

5 **Which term is defined as follows: a combination of substances not chemically bonded?**

- A atom
- B element
- C compound
- D mixture

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6 **Which of the following is an example of covalent bond?**

A sodium chloride held by opposite charges

B the oxygen hydrogen bond in water

C the bonds inside a water molecule

D the attraction between water molecules

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7 **Which statement best describes Avogadro's number?**

A a mole counts particles, not mass directly

B molar mass is found from atomic masses

C stoichiometry uses mole ratios from balanced equations

D Avogadro's number links moles to particle counts

Understand It with Pulse | Flag a mistake

8 **Which statement best describes Boyle's law?**

A gas pressure comes from particle collisions with walls

B Charles's law uses kelvin temperatures

C real gases behave most ideally at high temperature and low pressure

D Boyle's law is an inverse relationship

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9 **Which of the following is an example of compound?**

A a single oxygen atom

B gold, made only of gold atoms

C salt water before the salt dissolves fully

D water, made of hydrogen and oxygen

Understand It with Pulse | Flag a mistake

10 **Which of the following is an example of polar bond?**

A sodium chloride held by opposite charges

B the bonds inside a water molecule

C the attraction between water molecules

D the oxygen hydrogen bond in water

Understand It with Pulse | Flag a mistake

11 **Which statement best describes molar mass?**

A molar mass is found from atomic masses

B a mole counts particles, not mass directly

C Avogadro's number links moles to particle counts

D stoichiometry uses mole ratios from balanced equations

Understand It with Pulse | Flag a mistake

12 **Which statement best describes Charles's law?**

- A Charles's law uses kelvin temperatures
- B gas pressure comes from particle collisions with walls
- C Boyle's law is an inverse relationship
- D real gases behave most ideally at high temperature and low pressure

Understand It with Pulse | Flag a mistake

13 **Which of the following is an example of mixture?**

- A salt water before the salt dissolves fully
- B a single oxygen atom
- C gold, made only of gold atoms
- D water, made of hydrogen and oxygen

Understand It with Pulse | Flag a mistake

14 **Which of the following is an example of hydrogen bond?**

A the attraction between water molecules

B sodium chloride held by opposite charges

C the bonds inside a water molecule

D the oxygen hydrogen bond in water

Understand It with Pulse | Flag a mistake

15 **Which statement best describes stoichiometry?**

A a mole counts particles, not mass directly

B stoichiometry uses mole ratios from balanced equations

C Avogadro's number links moles to particle counts

D molar mass is found from atomic masses

Understand It with Pulse | Flag a mistake

16 **Which statement best describes ideal gas?**

A gas pressure comes from particle collisions with walls

B real gases behave most ideally at high temperature and low pressure

C Boyle's law is an inverse relationship

D Charles's law uses kelvin temperatures

Understand It with Pulse | Flag a mistake

17 **Which statement best describes compound?**

A a compound has properties unlike its elements

B atoms combine to form molecules and compounds

C elements are listed on the periodic table

D mixtures can be separated by physical means

Understand It with Pulse | Flag a mistake

18 **Which of the following is an example of chemical reaction?**

A sodium chloride held by opposite charges

B burning wood to form ash and gases

C the bonds inside a water molecule

D the oxygen hydrogen bond in water

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19 **Which term is defined as follows: a reactant left over after a reaction?**

A mole

B excess reactant

C Avogadro's number

D molar mass

Understand It with Pulse | Flag a mistake

20 **Which term is defined as follows: the temperature at which particle motion is lowest?**

- A pressure
- B absolute zero
- C Boyle's law
- D Charles's law

Understand It with Pulse | Flag a mistake

21 **Which of the following is an example of neutron?**

- A a single oxygen atom
- B gold, made only of gold atoms
- C a particle with no charge in an atom
- D water, made of hydrogen and oxygen

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22 **Which of the following is an example of reactant?**

A sodium chloride held by opposite charges

B the bonds inside a water molecule

C hydrogen gas before it reacts with oxygen

D the oxygen hydrogen bond in water

Understand It with Pulse | Flag a mistake

23 **Which term is defined as follows: actual yield divided by theoretical yield, times 100?**

A mole

B Avogadro's number

C percent yield

D molar mass

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24 **Which term is defined as follows: heat content of a system at constant pressure?**

A pressure

B Boyle's law

C enthalpy

D Charles's law

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25 **Which statement best describes electron?**

A electrons take part in bonding and reactions

B atoms combine to form molecules and compounds

C elements are listed on the periodic table

D a compound has properties unlike its elements

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26 **Which statement best describes product?**

A products appear on the right of an equation

B ionic compounds form crystal lattices

C covalent bonds hold atoms together in molecules

D polar bonds arise from differences in electronegativity

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27 **Which term is defined as follows: a uniform mixture with a solute dissolved in a solvent?**

A mole

B Avogadro's number

C molar mass

D solution

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28 **Which term is defined as follows: a reaction that releases heat?**

A pressure

B Boyle's law

C Charles's law

D exothermic reaction

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29 **Which term is defined as follows: a positively charged particle in a nucleus?**

A proton

B atom

C element

D compound

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30 **Which statement best describes synthesis reaction?**

A ionic compounds form crystal lattices

B synthesis builds a larger substance from smaller ones

C covalent bonds hold atoms together in molecules

D polar bonds arise from differences in electronegativity

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31 **Which of the following is an example of solute?**

A a dozen eggs as a counting unit

B salt in salt water

C the number of atoms in 12 grams of carbon 12

D 18 grams per mole for water

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32 **Which of the following is an example of endothermic reaction?**

- A a tire pressing on the road
- B ice melting in a warm room
- C squeezing a balloon raises its pressure
- D a balloon expanding when heated

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33 **Which statement best describes periodic table?**

- A atoms combine to form molecules and compounds
- B elements are listed on the periodic table
- C the periodic table arranges elements by atomic number
- D a compound has properties unlike its elements

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34 **Which statement best describes decomposition reaction?**

- A ionic compounds form crystal lattices
- B covalent bonds hold atoms together in molecules
- C decomposition splits a compound apart
- D polar bonds arise from differences in electronegativity

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35 **Which of the following is an example of solvent?**

- A a dozen eggs as a counting unit
- B the number of atoms in 12 grams of carbon 12
- C water in salt water
- D 18 grams per mole for water

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36 **Which of the following is an example of equilibrium?**

A a tire pressing on the road

B squeezing a balloon raises its pressure

C a tug of war with no movement

D a balloon expanding when heated

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37 **Which statement best describes metal?**

A atoms combine to form molecules and compounds

B elements are listed on the periodic table

C a compound has properties unlike its elements

D metals tend to lose electrons in reactions

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38 **Which statement best describes single replacement?**

A ionic compounds form crystal lattices

B covalent bonds hold atoms together in molecules

C polar bonds arise from differences in electronegativity

D single replacement follows an activity series

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39 **Which of the following is an example of concentration?**

A a dozen eggs as a counting unit

B the number of atoms in 12 grams of carbon 12

C 18 grams per mole for water

D moles per liter

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40 **Which of the following is an example of Le Chatelier's principle?**

A a tire pressing on the road

B squeezing a balloon raises its pressure

C a balloon expanding when heated

D adding reactant shifts a reaction toward products

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41 **Which statement best describes nonmetal?**

A nonmetals tend to gain electrons in reactions

B atoms combine to form molecules and compounds

C elements are listed on the periodic table

D a compound has properties unlike its elements

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42 **Which statement best describes double replacement?**

A double replacement often produces a solid, gas, or water

B ionic compounds form crystal lattices

C covalent bonds hold atoms together in molecules

D polar bonds arise from differences in electronegativity

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43 **Which statement best describes dilution?**

A a mole counts particles, not mass directly

B dilution increases volume while moles stay the same

C Avogadro's number links moles to particle counts

D molar mass is found from atomic masses

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44 **Which statement best describes acid?**

A gas pressure comes from particle collisions with walls

B acids have pH below 7 in water

C Boyle's law is an inverse relationship

D Charles's law uses kelvin temperatures

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45 **Which statement best describes element?**

A atoms combine to form molecules and compounds

B a compound has properties unlike its elements

C mixtures can be separated by physical means

D elements are listed on the periodic table

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46 **Which statement best describes precipitate?**

A ionic compounds form crystal lattices

B a precipitate signals that a reaction occurred

C covalent bonds hold atoms together in molecules

D polar bonds arise from differences in electronegativity

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47 **Which statement best describes balanced equation?**

A a mole counts particles, not mass directly

B Avogadro's number links moles to particle counts

C a balanced equation conserves atoms

D molar mass is found from atomic masses

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48 **Which statement best describes base?**

- A gas pressure comes from particle collisions with walls
- B Boyle's law is an inverse relationship
- C bases have pH above 7 in water
- D Charles's law uses kelvin temperatures

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49 **Which statement best describes molecule?**

- A atoms combine to form molecules and compounds
- B elements are listed on the periodic table
- C molecules are the units of molecular compounds
- D a compound has properties unlike its elements

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50 **Which statement best describes catalyst?**

A ionic compounds form crystal lattices

B covalent bonds hold atoms together in molecules

C a catalyst lowers activation energy

D polar bonds arise from differences in electronegativity

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ANSWER KEY

1.	A	**18.**	B	**35.**	C
2.	A	**19.**	B	**36.**	C
3.	B	**20.**	B	**37.**	D
4.	B	**21.**	C	**38.**	D
5.	D	**22.**	C	**39.**	D
6.	C	**23.**	C	**40.**	D
7.	D	**24.**	C	**41.**	A
8.	D	**25.**	A	**42.**	A
9.	D	**26.**	A	**43.**	B
10.	D	**27.**	D	**44.**	B
11.	A	**28.**	D	**45.**	D
12.	A	**29.**	A	**46.**	B
13.	A	**30.**	B	**47.**	C
14.	A	**31.**	B	**48.**	C
15.	B	**32.**	B	**49.**	C
16.	B	**33.**	C	**50.**	C
17.	A	**34.**	C		

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