MICHIGAN CHEMISTRY · UNIT 1 · 10 QUESTIONS

Matter and Atomic Structure — Practice Test

This practice test was made by AI and there may be mistakes. If a question or answer looks wrong, use the Flag a mistake button to let Brett know so he can fix it.

1 **The atomic number of an element is the number of**

- A protons
- B neutrons
- C electrons + neutrons
- D protons + neutrons

Understand It with Pulse | Flag a mistake

2 **Sand can be separated from water by**

- A evaporation only
- B filtration
- C condensation
- D freezing only

Understand It with Pulse | Flag a mistake

3 **A sample has mass 120 g and density 3 g per cm³. Its volume, in cm³, is _____**

ANSWER _____

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4 **The maximum number of electrons in the first energy level is**

A 2

B 8

C 18

D 32

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5 **Metals are typically**

A dull and brittle

B poor conductors of heat

C gases at room temperature

D good conductors of electricity

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6 **Using whole number atomic masses (C = 12, O = 16), the molar mass of carbon dioxide, in grams per mole, is ____**

ANSWER _____

Understand It with Pulse | Flag a mistake

7 **Carbon-12 and carbon-14 differ in their number of**

A protons

B neutrons

C electrons

D energy levels

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8 **Which of these is a homogeneous mixture**

A sand and water

B oil and water

C salt water

D iron filings and sand

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9 **An atom has mass number 27 and 14 neutrons. The number of protons in the atom is ____**

ANSWER _____

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10 **Elements in the periodic table are arranged in order of increasing**

A atomic mass

B atomic number

C density

D reactivity

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ANSWER KEY

1. A
2. B
3. 40
4. A
5. D
6. 44
7. B
8. C
9. 13
10. B

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