CHEMICAL BONDING, MOLECULAR SHAPE (VSEPR THEORY), INTERMOLECULAR FORCES, AND PROPERTIES - PRACTICE TEST | NEW BRUNSWICK CHEMISTRY 111-112 · UNIT B · 5 QUESTIONS

Chemical Bonding, Molecular Shape (VSEPR Theory), Intermolecular Forces, and Properties — Practice Test

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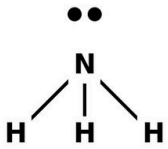
1

The Lewis diagram shows ammonia, with nitrogen bonded to three hydrogen atoms and a lone pair on nitrogen. The shape of the ammonia molecule is



Question 1

The Lewis diagram shows ammonia, with nitrogen bonded to three hydrogen atoms and a lone pair on nitrogen. The shape of the ammonia molecule is

A) linear

B) bent

C) trigonal pyramidal

D) tetrahedral

A linear

B bent

C trigonal pyramidal

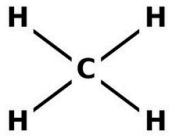
D tetrahedral

Understand It with Pulse

Flag a mistake

2

The Lewis diagram shows methane, with carbon bonded to four hydrogen atoms. The shape of the methane molecule is



Question 2

The Lewis diagram shows methane, with carbon bonded to four hydrogen atoms. The shape of the methane molecule is

A) linear

B) bent

C) trigonal planar

D) tetrahedral

A linear

B bent

C trigonal planar

D tetrahedral

Understand It with Pulse

Flag a mistake

3 **Neatly folded shirts and pairs of trousers in stacks. A lone pair on a central atom pushes bonding pairs**



A farther apart equally

B into a line always

C closer together

D out of the molecule

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4 **Sodium chloride is held together by**

A covalent bonds

B hydrogen bonds

C metallic bonds only

D ionic bonds

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5 **Diamond is hard because carbon atoms form a**

A strong covalent network

B weak ionic lattice

C gas

D liquid

Understand It with Pulse | Flag a mistake

ANSWER KEY

1. C
2. D
3. C
4. D
5. A

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