NEW HAMPSHIRE CHEMISTRY · UNIT 3 · 10 QUESTIONS

Stoichiometry and Solutions — Practice Test

This practice test was made by AI and there may be mistakes. If a question or answer looks wrong, use the Flag a mistake button to let Brett know so he can fix it.

1 **The molarity, in moles per liter, of a solution containing 0.5 mol of solute in 0.25 L is _____**

ANSWER _____

Understand It with Pulse | Flag a mistake

2 **A balanced chemical equation conserves**

A. mass

B. volume only

C. temperature

D. pressure

Understand It with Pulse | Flag a mistake

3 **The percent by mass of hydrogen in water, to one decimal place, is _____**

ANSWER _____

Understand It with Pulse | Flag a mistake

4 **After a reaction is complete, the excess reactant**



- **A** is completely used up
- **B** becomes a catalyst
- **C** disappears from the equation
- **D** remains in some amount

Understand It with Pulse | **Flag a mistake**

5 **In the reaction $CH_4 + 2O_2 \rightarrow CO_2 + 2H_2O$,**
the moles of carbon dioxide formed from 2 mol of methane is ____

ANSWER _____

Understand It with Pulse | Flag a mistake

6 **In the reaction $CaCO_3 \rightarrow CaO + CO_2$,**
the mass, in grams, of calcium oxide formed from 1 mol of calcium carbonate is ____

ANSWER _____

Understand It with Pulse | Flag a mistake

7 **In an aqueous solution, the solvent is**

A the solute

B water

C a gas

D a precipitate

Understand It with Pulse | Flag a mistake

8 **At standard temperature and pressure, the volume, in liters, of 2 mol of gas is ____**

ANSWER

Understand It with Pulse | Flag a mistake

9 **A reaction has a theoretical yield of 50 g and an actual yield of 40 g. The percent yield is**



A 125%

B 20%

C 80%

D 90%

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10 **The number of moles in 100 g of calcium carbonate is _____**

ANSWER _____

Understand It with Pulse | Flag a mistake

ANSWER KEY

1. 2
2. A
3. 11.1
4. D
5. 2
6. 56
7. B
8. 44.8
9. C
10. 1

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