NEWFOUNDLAND AND LABRADOR CHEMISTRY 2202 · UNIT A

# Stoichiometry — Practice Test

**Newfoundland and Labrador Chemistry 2202 • 10 questions**

This practice test was made by AI and there may be mistakes. If a question or answer looks wrong, use the Flag a mistake button to let Brett know so he can fix it.

**1** **Think about the property. An empirical formula gives**

- **A** the exact number of atoms in a molecule
- **B** the simplest whole number ratio of atoms
- **C** the molar mass
- **D** the percent composition only

Understand It with Pulse | Flag a mistake

**2** **In $2H_2 + O_2 \rightarrow 2H_2O$ the mass, in grams, of water formed from 4 mol of oxygen is**

ANSWER _____

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3 **A solution contains 4 mol of solute in 8 L. Its concentration, in moles per litre, is**

ANSWER _____

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4 **A reaction has theoretical yield 100 g and actual yield 80 g. The percent yield is**

ANSWER _____

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5 **Review the property. The molar mass of water, $H_2O$ in grams per mole, is**

ANSWER _____

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6 **The molar mass of ammonia, $NH_3$ in grams per mole, is**

ANSWER _____

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**7** **The amount, in moles, in 64 g of a substance with molar mass 32 g per mole is**

ANSWER _____

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**8** **The mass, in grams, of 3 mol of a substance with molar mass 18 g per mole is**

ANSWER _____

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**9** **Consider the statement in general. One mole of any substance contains about**

A $6.02 \times 10^{22}$ particles

B $6.02 \times 10^{23}$ particles

C $6.02 \times 10^{24}$ particles

D $6.02 \times 10^{20}$ particles

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10 **In the reaction $2H_2 + O_2 \rightarrow 2H_2O$ the moles of oxygen needed to react with 2 mol of hydrogen is**

ANSWER

Understand It with Pulse | Flag a mistake

# ANSWER KEY

| | | | | | |
|---|---|---|---|---|---|
| **1.** | B | **5.** | 18 | **9.** | B |
| **2.** | 144 | **6.** | 17 | **10.** | 1 |
| **3.** | 0.50 | **7.** | 2 | | |
| **4.** | 80 | **8.** | 54 | | |

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