NEWFOUNDLAND AND LABRADOR CHEMISTRY 2202 · UNIT B · 5 QUESTIONS

From Structures to Properties — Practice Test

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1 **Think about the idea. Water is a polar molecule mainly because**

A it is bent and oxygen attracts electrons more strongly

B it is linear

C it has no lone pairs

D hydrogen is more electronegative than oxygen

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2 **Hofmann voltameter water electrolysis apparatus on a laboratory bench: glass tubes filled with clear liquid with small gas bubbles rising from two electrodes, wires connected to a battery, bright daylight, no people, no text. Recall the idea. The strongest intermolecular force in liquid water is**



A London dispersion only

B hydrogen bonding

C dipole dipole only

D ionic bonding

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3 **Apply the idea. Ionic compounds in the molten state generally**

A never conduct electricity

B are gases at room temperature

C conduct electricity

D are soft and low melting

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4 **Consider the idea in general. Metallic bonding is best described as**

A sharing one pair of electrons

B transfer of protons

C hydrogen bonding between metals

D positive ions in a sea of delocalised electrons

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5

Review the idea. Across a period from left to right, electronegativity generally

A increases

B decreases

C stays the same

D becomes zero

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ANSWER KEY

1. A
2. B
3. C
4. D
5. A

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