NEWFOUNDLAND AND LABRADOR CHEMISTRY 3202 · UNIT C

Thermochemistry — Practice Test

Newfoundland and Labrador Chemistry 3202 • 10 questions

This practice test was made by AI and there may be mistakes. If a question or answer looks wrong, use the Flag a mistake button to let Brett know so he can fix it.

1 **Standard conditions for enthalpy values are usually 25 degrees and**

A 0 kilopascals

B 100 kilopascals

C 1000 kilopascals

D 1 kilopascal

Understand It with Pulse | Flag a mistake

2 **The enthalpy change of combustion is measured when one mole of a substance burns**

A partly in air

B without oxygen

C completely in oxygen

D in water

Understand It with Pulse | Flag a mistake

3 **A spontaneous process can still be slow if its**

A entropy is low

B enthalpy is negative

C temperature is high

D activation energy is high

Understand It with Pulse | Flag a mistake

4 **Freezing water decreases entropy because liquid water becomes a more**

A ordered solid

B disordered gas

C concentrated liquid

D heated solid

Understand It with Pulse | Flag a mistake

5 **Water mass 400 grams, specific heat capacity 4.19, temperature rise 2 degrees. The heat in joules is**

ANSWER _____

Understand It with Pulse | Flag a mistake

6 **A reaction releases 9000 joules for 3 moles. The molar enthalpy magnitude in joules per mole is**

ANSWER _____

Understand It with Pulse | Flag a mistake

7 **Enthalpy of products is 100 kilojoules and reactants is 350 kilojoules. The enthalpy change in kilojoules is**

ANSWER _____

Understand It with Pulse | Flag a mistake

8 **Bond energies give absorbed 600 and released 900 kilojoules. The enthalpy change in kilojoules is**

ANSWER _____

Understand It with Pulse | Flag a mistake

9 **Temperature rises from 18 degrees to 28 degrees. The change in degrees is**

ANSWER _____

Understand It with Pulse | Flag a mistake

10 **A fuel releases 30000 joules per gram. The heat from 4 grams in joules is**

ANSWER _____

Understand It with Pulse | Flag a mistake

ANSWER KEY

1.	B	**5.**	3352.0000000000005	**9.**	10
2.	C	**6.**	3000	**10.**	120000
3.	D	**7.**	-250		
4.	A	**8.**	-300		

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