NOVA SCOTIA ADVANCED CHEMISTRY 11 · UNIT A

Stoichiometry — Practice Test

Nova Scotia Advanced Chemistry 11 • 25 questions

This practice test was made by AI and there may be mistakes. If a question or answer looks wrong, use the Flag a mistake button to let Brett know so he can fix it.

1 **A reaction needs equal moles of two reactants. 4 moles of one and 9 moles of the other are mixed. The limiting amount, in moles, is**

ANSWER _____

Understand It with Pulse | Flag a mistake

2 **An advanced preparation has theoretical yield 120 grams and actual yield 96 grams. The percent yield is**

ANSWER _____

Understand It with Pulse | Flag a mistake

3 **An advanced sample of mass 75 grams has molar mass 25 grams per mole. The amount, in moles, is**

ANSWER _____

Understand It with Pulse | Flag a mistake

4 **A preparation has theoretical yield 160 grams and actual yield 128 grams. The percent yield is**

ANSWER _____

Understand It with Pulse | Flag a mistake

5 **A balanced equation for a reaction in solution must also balance**

A colour

B temperature

C pressure

D charge

Understand It with Pulse | Flag a mistake

6 **A burette is read to the nearest**

- A 0.05 millilitres in school work
- B 1 millilitre only
- C 10 millilitres
- D 0.5 litres

Understand It with Pulse | Flag a mistake

7 **A gas syringe measures the volume of gas**

- A dissolved in water
- B produced in a reaction
- C in a solid
- D absorbed by glass

Understand It with Pulse | Flag a mistake

8 **A hydrated salt contains water molecules as part of its**

A crystal structure

B colour only

C mass only

D solution

Understand It with Pulse | Flag a mistake

9 **A laboratory investigation report must state the method clearly enough that another student could**

A repeat the work

B guess the result

C skip safety steps

D change the aim

Understand It with Pulse | Flag a mistake

10 **A pipette delivers a volume that is**

- A estimated by eye
- B fixed and accurate
- C variable each time
- D measured by mass only

Understand It with Pulse | Flag a mistake

11 **A reaction uses 6 moles of reactant to make 12 moles of product. Starting with 3 moles of reactant, the product, in moles, is**

ANSWER _____

Understand It with Pulse | Flag a mistake

12 **A solution contains 4 moles in 2 litres. The concentration, in moles per litre, is**

ANSWER _____

Understand It with Pulse | Flag a mistake

13 **A solution of concentration 3 moles per litre has volume 4 litres. The amount, in moles, is**

ANSWER _____

Understand It with Pulse | Flag a mistake

14 **A standard solution has a concentration known**

A approximately only

B accurately

C by colour

D by mass only

Understand It with Pulse | Flag a mistake

15 **A standard solution is often made from a primary standard that is**

A coloured always

B pure and stable

C a gas only

D unstable

Understand It with Pulse | Flag a mistake

16 **A titre value far from the others should be checked for**

A automatic deletion

B inclusion without comment

C a procedural error before any decision

D rounding only

Understand It with Pulse | Flag a mistake

17 **A volumetric flask is used to make a solution of known**

- A volume and concentration
- B mass only
- C colour
- D temperature

Understand It with Pulse | Flag a mistake

18 **An advanced titration sample of mass 40 grams has molar mass 20 grams per mole. 2 moles of product form per mole of sample. The product amount, in moles, is**

ANSWER _____

Understand It with Pulse | Flag a mistake

19 **A reaction needs equal moles of two reactants. 3 moles of one and 8 moles of the other are mixed. The limiting amount, in moles, is**

ANSWER _____

Understand It with Pulse | Flag a mistake

20 **An advanced preparation has theoretical yield 200 grams and actual yield 170 grams. The percent yield is**

ANSWER _____

Understand It with Pulse | Flag a mistake

21 **An advanced sample of mass 96 grams has molar mass 48 grams per mole. The amount, in moles, is**

ANSWER _____

Understand It with Pulse | Flag a mistake

22 **A preparation has theoretical yield 250 grams and actual yield 200 grams. The percent yield is**

ANSWER _____

Understand It with Pulse | Flag a mistake

23 **A reaction needs equal moles of two reactants. 5 moles of one and 10 moles of the other are mixed. The limiting amount, in moles, is**

ANSWER _____

Understand It with Pulse | Flag a mistake

24 **An advanced preparation has theoretical yield 90 grams and actual yield 72 grams. The percent yield is**

ANSWER _____

Understand It with Pulse | Flag a mistake

25 **An advanced sample of mass 132 grams has molar mass 44 grams per mole. The amount, in moles, is**

ANSWER _____

Understand It with Pulse | Flag a mistake

ANSWER KEY

1.	4	**10.**	B	**19.**	3
2.	80	**11.**	6	**20.**	85
3.	3	**12.**	2	**21.**	2
4.	80	**13.**	12	**22.**	80
5.	D	**14.**	B	**23.**	5
6.	A	**15.**	B	**24.**	80
7.	B	**16.**	C	**25.**	3
8.	A	**17.**	A		
9.	A	**18.**	4		

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