NOVA SCOTIA ADVANCED CHEMISTRY 11 · UNIT B

From Structures to Properties — Practice Test

Nova Scotia Advanced Chemistry 11 • 10 questions

This practice test was made by AI and there may be mistakes. If a question or answer looks wrong, use the Flag a mistake button to let Brett know so he can fix it.

1 **A molecule with two bonding pairs and no lone pairs is**

- A linear
- B bent
- C tetrahedral
- D pyramidal

Understand It with Pulse | Flag a mistake

2 **Sulfur hexafluoride has a shape described as**

A tetrahedral

B octahedral

C linear

D bent

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3 **A polar bond forms between atoms with different**

A masses only

B colours

C electronegativities

D sizes only

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4 **Hydrogen bonding explains why water has a boiling point that is**

A lower than expected

B zero

C equal to methane

D higher than expected for its mass

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5 **A network covalent solid is usually**

A hard with a high melting point

B soft with a low melting point

C a gas at room temperature

D soluble in water

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6 **Simple molecular substances are usually poor conductors because they have no**

A covalent bonds

B free charged particles

C atoms

D electrons

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7 **The first shell of an atom holds at most**

A 8 electrons

B 18 electrons

C 2 electrons

D 1 electron

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8 **The second shell of an atom holds at most**

- A 2 electrons
- B 18 electrons
- C 32 electrons
- D 8 electrons

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9 **An isotope has the same atomic number but a different**

- A mass number
- B element name
- C charge always
- D electron count in a neutral atom

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10 **A dative covalent bond forms when both shared electrons come from**

A two atoms equally

B one atom

C a proton

D a neutron

Understand It with Pulse | Flag a mistake

ANSWER KEY

1.	A	**5.**	A	**9.**	A
2.	B	**6.**	B	**10.**	B
3.	C	**7.**	C		
4.	D	**8.**	D		

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