NOVA SCOTIA ADVANCED CHEMISTRY 12 · UNIT B

Equilibrium: Acids and Bases in Chemical Changes — Practice Test

Nova Scotia Advanced Chemistry 12 • 10 questions

This practice test was made by AI and there may be mistakes. If a question or answer looks wrong, use the Flag a mistake button to let Brett know so he can fix it.

1 **A solution has pOH 0. The hydroxide ion concentration, in moles per litre, is**

ANSWER _____

Understand It with Pulse | Flag a mistake

2 **At equilibrium a product concentration is 0.8 moles per litre and the reactant concentration is 0.4 moles per litre. For a reaction making 2 product particles from 1 reactant particle the equilibrium constant is**

ANSWER _____

Understand It with Pulse | Flag a mistake

3 **At equilibrium a product concentration is 1.0 moles per litre and the reactant concentration is 0.5 moles per litre. For a reaction making 2 product particles from 1 reactant particle the equilibrium constant is**

ANSWER _____

Understand It with Pulse | Flag a mistake

4 **50 millilitres of acid neutralise 20 millilitres of 0.1 moles per litre base in a 1 to 1 reaction. The acid concentration, in moles per litre, is**

ANSWER _____

Understand It with Pulse | Flag a mistake

5 **20 millilitres of acid neutralise 40 millilitres of 0.1 moles per litre base in a 1 to 1 reaction. The acid concentration, in moles per litre, is**

ANSWER _____

Understand It with Pulse | Flag a mistake

6 **A weak acid of concentration 0.5 moles per litre has hydrogen ion concentration 0.05 moles per litre. The acid dissociation constant is about**

ANSWER _____

Understand It with Pulse | Flag a mistake

7 **A weak acid of concentration 1.0 moles per litre has hydrogen ion concentration 0.1 moles per litre. The acid dissociation constant is about**

ANSWER _____

Understand It with Pulse | Flag a mistake

8 **A solution with pH 2 is classified as**

A strongly acidic

B weakly acidic

C neutral

D basic

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9 **Adding sodium chloride to a saturated solution of silver chloride lowers how much silver chloride can dissolve. The added chloride is called a**

A spectator ion

B common ion

C catalyst

D buffer

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10 **Titration with a pH meter finds the equivalence point from a sharp**

A colour change

B mass change

C pH jump

D volume fall

Understand It with Pulse | Flag a mistake

ANSWER KEY

1.	1	**5.**	0.2	**9.**	B
2.	1.6	**6.**	0.005	**10.**	C
3.	2	**7.**	0.01		
4.	0.04	**8.**	A		

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