NORTHWEST TERRITORIES CHEMISTRY 12 (MCH--12) · UNIT B · 25 QUESTIONS

Chemical Equilibrium — Practice Test

This practice test was made by AI and there may be mistakes. If a question or answer looks wrong, use the Flag a mistake button to let Brett know so he can fix it.

1 **At dynamic equilibrium in a closed system**

- A all reactions stop
- B the forward and reverse reactions continue at equal rates
- C all reactants have been used up
- D only the forward reaction continues

Understand It with Pulse | Flag a mistake

2 **At equilibrium, the concentrations of reactants and products**

- A remain constant, though not necessarily equal
- B are always equal to each other
- C keep changing until the reaction stops
- D are always zero for the reactants

Understand It with Pulse | Flag a mistake

3 **The equilibrium law expression for $H_2(g) + I_2(g) \rightleftharpoons 2HI(g)$ is**

- **A** $K_eq = \frac{[HI]}{[H_2][I_2]}$
- **B** $K_eq = \frac{[H_2][I_2]}{[HI]^2}$
- **C** $K_eq = \frac{[HI]^2}{[H_2][I_2]}$
- **D** $K_eq = [HI]^2[H_2][I_2]$

Understand It with Pulse | Flag a mistake

4 **The equilibrium law expression for $N_2(g) + 3H_2(g) \rightleftharpoons 2NH_3(g)$ is**

- **A** $K_eq = \frac{[NH_3]}{[N_2][H_2]}$
- **B** $K_eq = \frac{[N_2][H_2]^3}{[NH_3]^2}$
- **C** $K_eq = \frac{[NH_3]^2}{[N_2][H_2]}$
- **D** $K_eq = \frac{[NH_3]^2}{[N_2][H_2]^3}$

Understand It with Pulse | Flag a mistake

5 **For $CaCO_3(s) \rightleftharpoons CaO(s) + CO_2(g)$, the equilibrium expression is**

- **A** $K_eq = [CaO][CO_2]$
- **B** $K_eq = \frac{[CaO][CO_2]}{[CaCO_3]}$
- **C** $K_eq = [CO_2]$
- **D** $K_eq = \frac{1}{[CO_2]}$

Understand It with Pulse | Flag a mistake

6 **At equilibrium, $[H_2] = 0.10$, $[I_2] = 0.10$, and $[HI] = 0.40$, all in mol/L, for $H_2 + I_2 \rightleftharpoons 2HI$. K_eq is**

- **A** 4
- **B** 16
- **C** 40
- **D** 160

Understand It with Pulse | Flag a mistake

7 At equilibrium, $[N_2] = 0.25$, $[H_2] = 0.50$, and $[NH_3] = 0.50$, all in mol/L, for $N_2 + 3H_2 \rightleftharpoons 2NH_3$. K_eq is

- A 8
- B 2
- C 4
- D 16

Understand It with Pulse | Flag a mistake

8 At equilibrium, $[CO] = 0.20$, $[H_2O] = 0.20$, $[CO_2] = 0.40$, and $[H_2] = 0.40$, all in mol/L, for $CO + H_2O \rightleftharpoons CO_2 + H_2$. The value of K_eq is

- A 1
- B 2
- C 8
- D 4

Understand It with Pulse | Flag a mistake

9 **A large K_eq value means that at equilibrium**

- A reactants are favoured over products
- B products are favoured over reactants
- C the reaction is very slow
- D no products have formed

Understand It with Pulse | Flag a mistake

10 **A very small K_eq value means that at equilibrium**

- A products are favoured over reactants
- B the forward and reverse rates are zero
- C reactants are favoured over products
- D the reaction cannot be reversed

Understand It with Pulse | Flag a mistake

11 For A ⇌ B, an ICE table starts with [A] = 1.00 mol/L and [B] = 0. At equilibrium [B] = 0.20 mol/L. The equilibrium concentration of A, in mol/L, is

- A 0.20
- B 0.40
- C 1.20
- D 0.80

Understand It with Pulse | Flag a mistake

12 For A ⇌ B, with [A] falling from 1.00 to 0.80 mol/L, and [B] starting at 0, the value of K_eq is

- A 0.25
- B 0.20
- C 0.80
- D 4.0

Understand It with Pulse | Flag a mistake

13 **For $H_2 + I_2 \rightleftharpoons 2HI$, both reactants start at 0.50 mol/L. At equilibrium [HI] = 0.60 mol/L. The equilibrium [H_2], in mol/L, is**

A 0.10

B 0.30

C 0.40

D 0.20

Understand It with Pulse | Flag a mistake

14 **For the system in the previous type of ICE problem, with [H_2] = 0.20, [I_2] = 0.20, and [HI] = 0.60, all in mol/L, K_eq is**

A 3

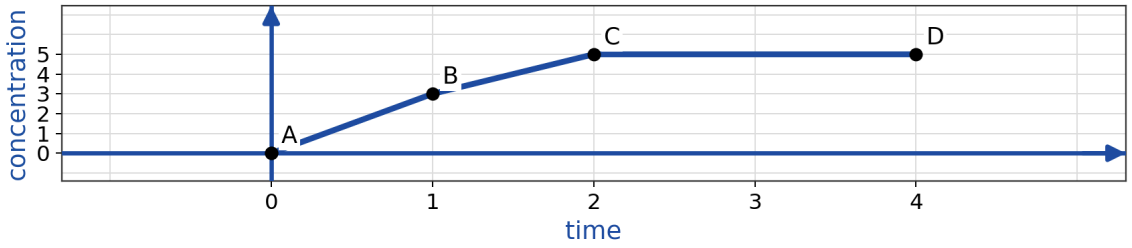
B 6

C 9

D 15

Understand It with Pulse | Flag a mistake

15 **The graph shows the concentration of a product against time. Points are (0, 0), (1, 3), (2, 5), and (4, 5). Equilibrium is first reached at about**



A 1 minute

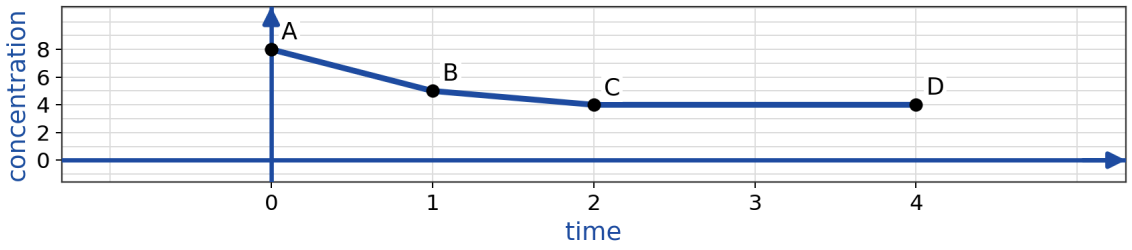
B 2 minutes

C 4 minutes

D 0 minutes

Understand It with Pulse | Flag a mistake

16 **The graph shows the concentration of a reactant against time. Points are (0, 8), (1, 5), (2, 4), and (4, 4). The flat section of the graph shows that**



- **A** equilibrium has been reached
- **B** the reaction has stopped completely
- **C** all reactant has been used up
- **D** the temperature has fallen to zero

Understand It with Pulse | Flag a mistake

17 **For $N_2 + 3H_2 \rightleftharpoons 2NH_3$ + heat, adding more N_2 at equilibrium will shift the equilibrium to the**

A right, forming more NH_3

B left, forming more N_2

C right, forming more H_2

D left, with no change in NH_3

Understand It with Pulse | Flag a mistake

18 **For $N_2 + 3H_2 \rightleftharpoons 2NH_3$ + heat, removing NH_3 at equilibrium will shift the equilibrium to the**

A left, to use up H_2

B right, to use up NH_3

C right, to replace the NH_3 removed

D left, with no change

Understand It with Pulse | Flag a mistake

19 **For the exothermic reaction N_2 + $3H_2 \rightleftharpoons 2NH_3$ + heat, raising the temperature will shift the equilibrium to the**

A right, favouring the products

B left, favouring the reactants

C left, favouring the products

D right, with no change in K_eq

Understand It with Pulse | Flag a mistake

20 **For the endothermic reaction N_2 + O_2 + heat $\rightleftharpoons 2NO$, raising the temperature will shift the equilibrium to the**

A left, favouring the reactants

B right, favouring the reactants

C left, with no change in K_eq

D right, favouring the products

Understand It with Pulse | Flag a mistake

21 **For $N_2 + 3H_2 \rightleftharpoons 2NH_3$, increasing the pressure by decreasing volume will shift the equilibrium toward the side with**

A more moles of gas, the reactant side

B fewer moles of gas, the reactant side

C fewer moles of gas, the product side

D more moles of gas, the product side

Understand It with Pulse | Flag a mistake

22 **For $H_2 + I_2 \rightleftharpoons 2HI$, changing the pressure will**

A not shift the equilibrium, since gas moles are equal on both sides

B shift it to the right

C shift it to the left

D change K_eq but not the position

Understand It with Pulse | Flag a mistake

23 **Adding a catalyst to a system at equilibrium will**

- A shift the equilibrium to the right
- B shift the equilibrium to the left
- C increase K_eq
- D not change the equilibrium position or K_eq

Understand It with Pulse | Flag a mistake

24 **Which change alters the value of K_eq for a reaction**

- A a change in concentration
- B a change in temperature
- C a change in pressure
- D adding a catalyst

Understand It with Pulse | Flag a mistake

25 **If Q is less than K_eq, the reaction will proceed to the**

A left, forming more reactants

B right, forming more products

C right, forming more reactants

D left, until Q falls to zero

Understand It with Pulse | Flag a mistake

ANSWER KEY

1. B
2. A
3. C
4. D
5. C
6. B
7. A
8. D
9. B
10. C
11. D
12. A
13. D
14. C
15. B
16. A
17. A
18. C
19. B
20. D
21. C
22. A
23. D
24. B
25. B

© 2026 Brett Robbins. All Rights Reserved.