OHIO CHEMISTRY · UNIT 3

Stoichiometry and Solutions — Practice Test

Ohio Chemistry • 10 questions

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1 **The molarity, in moles per liter, of a solution containing 0.5 mol of solute in 0.25 L is ____**

ANSWER _____

Understand It with Pulse | Flag a mistake

2 **A balanced chemical equation conserves**

A mass

B volume only

C temperature

D pressure

Understand It with Pulse | Flag a mistake

3 **The percent by mass of hydrogen in water, to one decimal place, is _____**

ANSWER _____

Understand It with Pulse | Flag a mistake

4 **After a reaction is complete, the excess reactant**

A is completely used up

B becomes a catalyst

C disappears from the equation

D remains in some amount

Understand It with Pulse | Flag a mistake

5 **In the reaction $CH_4 + 2O_2 \rightarrow CO_2 + 2H_2O$,**
the moles of carbon dioxide formed from 2 mol of methane is ____

ANSWER _____

Understand It with Pulse | Flag a mistake

6 **In the reaction $CaCO_3 \rightarrow CaO + CO_2$, the mass, in grams, of calcium oxide formed from 1 mol of calcium carbonate is ____**

ANSWER _____

Understand It with Pulse | Flag a mistake

7 **In an aqueous solution, the solvent is**

A the solute

B water

C a gas

D a precipitate

Understand It with Pulse | Flag a mistake

8 **At standard temperature and pressure, the volume, in liters, of 2 mol of gas is ____**

ANSWER _____

Understand It with Pulse | Flag a mistake

9 **A reaction has a theoretical yield of 50 g and an actual yield of 40 g. The percent yield is**

- **A** 125%
- **B** 20%
- **C** 80%
- **D** 90%

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10 **The number of moles in 100 g of calcium carbonate is ____**

ANSWER _____

Understand It with Pulse | Flag a mistake

ANSWER KEY

1.	2	**5.**	2	**9.**	C
2.	A	**6.**	56	**10.**	1
3.	11.1	**7.**	B		
4.	D	**8.**	44.8		

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