ONTARIO CHEMISTRY, GRADE 11, UNIVERSITY PREPARATION (SCH3U) · UNIT C · 5 QUESTIONS

Quantities in Chemical Reactions — Practice Test

This practice test was made by AI and there may be mistakes. If a question or answer looks wrong, use the Flag a mistake button to let Brett know so he can fix it.

1

Scuba diver swimming underwater with a silver oxygen tank. In the reaction 2 Mg plus O_2 gives 2 MgO, how many moles of magnesium oxide form from 4 moles of magnesium, with excess oxygen?



ANSWER

Understand It with Pulse

Flag a mistake

2 **The limiting reactant in a reaction is the reactant that**

A is present in excess

B is a catalyst

C has the largest molar mass

D is used up first and limits the amount of product

Understand It with Pulse | Flag a mistake

3 **Percent yield is calculated as**

A actual yield ÷ theoretical yield, × 100 percent

B actual yield × theoretical yield, × 100 percent

C theoretical yield ÷ actual yield, × 100 percent

D actual yield + theoretical yield

Understand It with Pulse | Flag a mistake

4 **A reaction has an actual yield of 40 grams and a theoretical yield of 50 grams. What is the percent yield?**

ANSWER _____

Understand It with Pulse | Flag a mistake

5 **A reaction has an actual yield of 36 grams and a theoretical yield of 60 grams. What is the percent yield?**

ANSWER _____

Understand It with Pulse | Flag a mistake

ANSWER KEY

1. 4
2. D
3. A
4. 80
5. 60

© 2026 Brett Robbins. All Rights Reserved.