ONTARIO CHEMISTRY, GRADE 12, UNIVERSITY PREPARATION (SCH4U) · UNIT D · 5 QUESTIONS

Chemical Systems, Equilibrium, and Electrochemistry — Practice Test

This practice test was made by AI and there may be mistakes. If a question or answer looks wrong, use the Flag a mistake button to let Brett know so he can fix it.

1 **The reaction quotient Q compared with K predicts direction: if Q is less than K, the reaction shifts toward**

A Products

B Neither

C The catalyst

D Reactants

Understand It with Pulse | Flag a mistake

2 **Pure solids and liquids are omitted from equilibrium expressions because their concentrations are**

- A Zero
- B Constant
- C Very large
- D Unknown

Understand It with Pulse | Flag a mistake

3 **For a gaseous equilibrium, Kp uses**

- A Volumes only
- B Molar concentrations
- C Masses
- D Partial pressures

Understand It with Pulse | Flag a mistake

4 **For the Haber process, increasing pressure favours ammonia formation because the product side has**

- A More gas moles
- B Fewer gas moles
- C Equal gas moles
- D No gas

Understand It with Pulse | Flag a mistake

5 **The Haber process uses a catalyst of**

- A Copper
- B Nickel
- C Platinum
- D Iron

Understand It with Pulse | Flag a mistake

ANSWER KEY

1. A
2. B
3. D
4. B
5. D

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