PRINCE EDWARD ISLAND CHEMISTRY 11 · UNIT B

Stoichiometry — Practice Test

Prince Edward Island Chemistry 11 • 10 questions

This practice test was made by AI and there may be mistakes. If a question or answer looks wrong, use the Flag a mistake button to let Brett know so he can fix it.

1 **What is the mass in grams of 2 moles of water, if the molar mass of water is 18 grams per mole**

ANSWER

Understand It with Pulse | Flag a mistake

2 **What is the mass in grams of 3 moles of carbon dioxide, if the molar mass of carbon dioxide is 44 grams per mole**

ANSWER

Understand It with Pulse | Flag a mistake

3 **The molar mass of water is about**

- A 10 grams per mole
- B 32 grams per mole
- C 18 grams per mole
- D 44 grams per mole

Understand It with Pulse | Flag a mistake

4 **The molecular formula of glucose is**

- A CH_2O
- B $C_2H_4O_2$
- C $C_6H_{12}O_6$
- D C_6H_6

Understand It with Pulse | Flag a mistake

5 **The theoretical yield is the maximum amount of product possible if the reaction goes to**

A equilibrium

B zero

C halfway

D completion

Understand It with Pulse | Flag a mistake

6 **To convert grams to moles, you**

A divide by the molar mass

B multiply by the molar mass

C add the molar mass

D subtract the molar mass

Understand It with Pulse | Flag a mistake

7 **To convert moles to grams, you**

A divide by the molar mass

B multiply by the molar mass

C add the molar mass

D subtract the molar mass

Understand It with Pulse | Flag a mistake

8 **What is the mass in grams of 4 moles of water, if the molar mass of water is 18 grams per mole**

ANSWER _____

Understand It with Pulse | Flag a mistake

9 **What is the mass in grams of 2 moles of carbon dioxide, if the molar mass of carbon dioxide is 44 grams per mole**

ANSWER _____

Understand It with Pulse | Flag a mistake

10 **What is the mass in grams of 5 moles of water, if the molar mass of water is 18 grams per mole**

ANSWER

Understand It with Pulse | Flag a mistake

ANSWER KEY

1.	36	**5.**	D	**9.**	88
2.	132	**6.**	A	**10.**	90
3.	C	**7.**	B		
4.	C	**8.**	72		

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