PRINCE EDWARD ISLAND CHEMISTRY 11 · FINAL EXAM

Final Exam — Practice Test

Prince Edward Island Chemistry 11 • 50 questions

This practice test was made by AI and there may be mistakes. If a question or answer looks wrong, use the Flag a mistake button to let Brett know so he can fix it.

1 **Sodium chloride is an example of an**

- A ionic compound
- B covalent compound
- C metallic element
- D mixture

Understand It with Pulse | Flag a mistake

2 **In the reaction $2H_2$ plus O_2 gives $2H_2O$, 2 moles of hydrogen react completely. How many moles of water are produced**

ANSWER _____

Understand It with Pulse | Flag a mistake

3 **An alkane with 1 carbon atoms has a molar mass in grams per mole of 12 times the carbon atoms plus 1 times the hydrogen atoms. What is its molar mass, using hydrogen atoms from the alkane formula**

ANSWER _____

Understand It with Pulse | Flag a mistake

4 **An experiment is repeated 5 times, and each trial records 10 observations. How many observations are recorded in total**

ANSWER _____

Understand It with Pulse | Flag a mistake

5 **An atom of sodium has atomic number 11 and mass number 23. How many neutrons does the atom have**

ANSWER _____

Understand It with Pulse | Flag a mistake

6 **A balanced chemical equation has the same number of each type of atom on**

- A both sides
- B the left only
- C the right only
- D neither side

Understand It with Pulse | Flag a mistake

7 **An alkane contains only**

- A single bonds
- B double bonds
- C triple bonds
- D ionic bonds

Understand It with Pulse | Flag a mistake

8 **A scientific inference is a conclusion based on**

A opinions only

B observations

C guesses only

D theories only

Understand It with Pulse | Flag a mistake

9 **Covalent bonds are formed by the sharing of**

A protons

B neutrons

C ions

D electrons

Understand It with Pulse | Flag a mistake

10 **Coefficients in a balanced equation show the**

A mass ratio only

B mole ratio of reactants and products

C volume ratio only

D temperature

Understand It with Pulse | Flag a mistake

11 **Benzene is an example of an**

A alkane

B aromatic hydrocarbon

C alkene only

D alcohol

Understand It with Pulse | Flag a mistake

12 **A scientific theory is a well tested explanation that**

A is a guess

B is a law

C cannot be tested

D explains many observations

Understand It with Pulse | Flag a mistake

13 **A neutral atom of carbon has atomic number 6. How many electrons does the neutral atom have**

ANSWER _____

Understand It with Pulse | Flag a mistake

14 **One mole of a substance contains the same number of particles as one mole of**

A any other substance

B water only

C carbon only

D one gram of it

Understand It with Pulse | Flag a mistake

15 **Crude oil is separated by fractional distillation based on differences in**

A melting point only

B density only

C boiling point

D colour

Understand It with Pulse | Flag a mistake

16 **Dalton proposed that matter is made of tiny indivisible particles called**

A atoms

B molecules only

C ions

D electrons

Understand It with Pulse | Flag a mistake

17 **Covalent molecular substances usually have melting points that are**

A very high

B zero

C low

D unpredictable

Understand It with Pulse | Flag a mistake

18 **Stoichiometry is the study of quantitative relationships in**

- A nuclear reactions only
- B physical changes only
- C solutions only
- D chemical reactions

Understand It with Pulse | Flag a mistake

19 **Isomers are compounds with the same molecular formula but different**

- A molar masses
- B numbers of atoms
- C structures
- D elements

Understand It with Pulse | Flag a mistake

20 **Rutherford proposed that an atom has a small dense**

A electron cloud only

B plum pudding

C orbit only

D nucleus

Understand It with Pulse | Flag a mistake

21 **A neutral atom of chlorine has atomic number 17. How many electrons does the neutral atom have**

ANSWER _____

Understand It with Pulse | Flag a mistake

22 **The limiting reactant is the reactant that is**

- A left over
- B in excess
- C used up first
- D a catalyst

Understand It with Pulse | Flag a mistake

23 **An alkane with 2 carbon atoms has a molar mass in grams per mole of 12 times the carbon atoms plus 1 times the hydrogen atoms. What is its molar mass, using hydrogen atoms from the alkane formula**

ANSWER _____

Understand It with Pulse | Flag a mistake

24 **The Bohr atomic model dates to about 1913, and the Chadwick model dates to about 1932. How many years apart are the two models**

ANSWER _____

Understand It with Pulse | Flag a mistake

25 **An atom has 6 protons and 6 neutrons. What is the mass number of the atom**

ANSWER _____

Understand It with Pulse | Flag a mistake

26 **In the reaction $2H_2$ plus O_2 gives $2H_2O$, 4 moles of water are produced. How many moles of oxygen were used**

ANSWER _____

Understand It with Pulse | Flag a mistake

27 **An alkane has 3 carbon atoms. Using the formula C n H 2n plus 2, how many hydrogen atoms does the alkane have**

ANSWER _____

Understand It with Pulse | Flag a mistake

28 **An experiment is repeated 4 times, and each trial records 10 observations. How many observations are recorded in total**

ANSWER _____

Understand It with Pulse | Flag a mistake

29 **The valence electrons of an atom are the electrons in its**

A outermost shell

B innermost shell

C nucleus

D first shell only

Understand It with Pulse | Flag a mistake

30 **In the reaction $2H_2$ plus O_2 gives $2H_2O$, 6 moles of hydrogen react completely. How many moles of water are produced**

ANSWER _____

Understand It with Pulse | Flag a mistake

31 **An alkane has 4 carbon atoms. Using the formula C n H 2n plus 2, how many hydrogen atoms does the alkane have**

ANSWER _____

Understand It with Pulse | Flag a mistake

32 **A measured value is 90 and the accepted value is 100. What is the percent error, in percent**

ANSWER _____

Understand It with Pulse | Flag a mistake

33 **An atom has 17 protons and 18 neutrons. What is the mass number of the atom**

ANSWER _____

Understand It with Pulse | Flag a mistake

34 **In the reaction $2H_2$ plus O_2 gives $2H_2O$, 8 moles of water are produced. How many moles of oxygen were used**

ANSWER _____

Understand It with Pulse | Flag a mistake

35 **An alkane has 5 carbon atoms. Using the formula C n H 2n plus 2, how many hydrogen atoms does the alkane have**

ANSWER _____

Understand It with Pulse | Flag a mistake

36 **A measured value is 120 and the accepted value is 100. What is the percent error, in percent**

ANSWER _____

Understand It with Pulse | Flag a mistake

37 **Atomic radius generally decreases across a period because the nuclear charge**

A decreases

B stays constant

C increases

D becomes zero

Understand It with Pulse | Flag a mistake

38 **In the reaction $2H_2$ plus O_2 gives $2H_2O$, 10 moles of water are produced. How many moles of oxygen were used**

ANSWER _____

Understand It with Pulse | Flag a mistake

39 **An alkene has 6 carbon atoms. Using the formula C n H 2n, how many hydrogen atoms does the alkene have**

ANSWER _____

Understand It with Pulse | Flag a mistake

40 **A controlled experiment changes one variable while keeping other variables**

A changing

B constant

C ignored

D random always

Understand It with Pulse | Flag a mistake

41 **A neutral atom of calcium loses 2 electrons to form an ion. What is the charge of the ion, as a positive whole number**

ANSWER _____

Understand It with Pulse | Flag a mistake

42 **The molar mass of water is about**

A 10 grams per mole

B 32 grams per mole

C 18 grams per mole

D 44 grams per mole

Understand It with Pulse | Flag a mistake

43 **The functional group of an alcohol is the**

A carboxyl group

B hydroxyl group

C carbonyl group

D amino group

Understand It with Pulse | Flag a mistake

44 **The Rutherford atomic model dates to about 1911, and the Bohr model dates to about 1913. How many years apart are the two models**

ANSWER _____

Understand It with Pulse | Flag a mistake

45 **Electronegativity is a measure of an atom's ability to**

A lose protons

B attract bonding electrons

C gain neutrons

D form ions

Understand It with Pulse | Flag a mistake

46

To convert grams to moles, you

A divide by the molar mass

B multiply by the molar mass

C add the molar mass

D subtract the molar mass

Understand It with Pulse | Flag a mistake

47

Unsaturated hydrocarbons contain at least one double or

A triple bond

B single bond only

C ionic bond

D metallic bond

Understand It with Pulse | Flag a mistake

48 **Three repeated measurements give 15, 17, and 16. What is the mean of the three measurements, to two decimal places**

ANSWER _____

Understand It with Pulse | Flag a mistake

49 **A neutral atom of potassium loses 1 electrons to form an ion. What is the charge of the ion, as a positive whole number**

ANSWER _____

Understand It with Pulse | Flag a mistake

50 **What is the mass in grams of 2 moles of carbon dioxide, if the molar mass of carbon dioxide is 44 grams per mole**

ANSWER _____

Understand It with Pulse | Flag a mistake

ANSWER KEY

1.	A	**18.**	D	**35.**	12
2.	2	**19.**	C	**36.**	20.00
3.	16	**20.**	D	**37.**	C
4.	50	**21.**	17	**38.**	5
5.	12	**22.**	C	**39.**	12
6.	A	**23.**	30	**40.**	B
7.	A	**24.**	19	**41.**	2
8.	B	**25.**	12	**42.**	C
9.	D	**26.**	2	**43.**	B
10.	B	**27.**	8	**44.**	2
11.	B	**28.**	40	**45.**	B
12.	D	**29.**	A	**46.**	A
13.	6	**30.**	6	**47.**	A
14.	A	**31.**	10	**48.**	16.00
15.	C	**32.**	10.00	**49.**	1
16.	A	**33.**	35	**50.**	88
17.	C	**34.**	4		

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