QUEBEC CHEMISTRY, SECONDARY V · UNIT A

Gases — Practice Test

Quebec Chemistry, Secondary V • 10 questions

This practice test was made by AI and there may be mistakes. If a question or answer looks wrong, use the Flag a mistake button to let Brett know so he can fix it.

1 **Increasing the temperature of a gas increases the average kinetic energy of its**

- A container
- B colour
- C mass
- D particles

Understand It with Pulse | Flag a mistake

2 **Real gases behave most like ideal gases at high temperature and**

A high pressure

B low pressure

C low temperature

D high density

Understand It with Pulse | Flag a mistake

3 **Standard temperature and pressure means 0 degrees Celsius and a pressure of**

A 101.3 kilopascals

B 0 kilopascals

C 200 kilopascals

D 50 kilopascals

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4 **The combined gas law combines Boyle law, Charles law and**

A Dalton law

B Avogadro law only

C Hess law

D Gay Lussac law

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5 **The Kelvin temperature scale starts at absolute zero, which is -273 degrees**

A Celsius

B Fahrenheit

C Kelvin

D Rankine

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6 **The molar volume of an ideal gas at standard temperature and pressure is about**

- A 1 litre per mole
- B 100 litres per mole
- C 0 litres per mole
- D 22.4 litres per mole

Understand It with Pulse | Flag a mistake

7 **The partial pressure of a gas in a mixture depends on its mole**

- A colour
- B mass only
- C fraction
- D temperature only

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8 **The pressure of a gas is caused by particles colliding with the**

A floor only

B ceiling only

C walls of the container

D other gases only

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9 **To convert Celsius to Kelvin you**

A subtract 273

B add 273

C multiply by 273

D divide by 273

Understand It with Pulse | Flag a mistake

10 **When a gas expands against a constant external pressure it does**

A nothing

B melting

C cooling only

D work

Understand It with Pulse | Flag a mistake

ANSWER KEY

1.	D	**5.**	A	**9.**	B
2.	B	**6.**	D	**10.**	D
3.	A	**7.**	C		
4.	D	**8.**	C		

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