QUEBEC CHEMISTRY, SECONDARY V · UNIT D

Chemical Equilibrium — Practice Test

Quebec Chemistry, Secondary V • 10 questions

This practice test was made by AI and there may be mistakes. If a question or answer looks wrong, use the Flag a mistake button to let Brett know so he can fix it.

1 **If Kc is 1, reactants and products are present in comparable**

- A colours
- B masses only
- C temperatures
- D amounts

Understand It with Pulse | Flag a mistake

2 **Increasing pressure shifts a gaseous equilibrium toward the side with fewer moles of**

A solid

B liquid

C gas

D catalyst

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3 **Increasing temperature shifts an endothermic equilibrium toward the**

A products

B reactants

C catalyst

D solids

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4 **Increasing temperature shifts an exothermic equilibrium toward the**

A products

B catalyst

C solids

D reactants

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5 **Le Chatelier principle states that a system at equilibrium shifts to counteract an imposed**

A catalyst

B colour

C container

D change

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6 **Pure solids and pure liquids are omitted from an equilibrium expression because their concentrations are**

A constant

B zero

C infinite

D unknown

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7 **Removing product from an equilibrium system shifts it toward the**

A reactants

B products

C catalyst

D solids

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8 **The Contact process makes sulfuric acid and uses vanadium pentoxide as a**

A reactant

B product

C catalyst

D fuel

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9 **The equilibrium constant changes only when the**

A pressure changes

B catalyst changes

C temperature changes

D volume changes

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10 **The Haber process makes ammonia from nitrogen and**

A oxygen

B hydrogen

C carbon dioxide

D methane

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ANSWER KEY

1.	D	**5.**	D	**9.**	C
2.	C	**6.**	A	**10.**	B
3.	A	**7.**	B		
4.	D	**8.**	C		

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