SASKATCHEWAN PHYSICAL SCIENCE 20 · UNIT C

The Mole and Stoichiometry — Practice Test

Saskatchewan Physical Science 20 • 10 questions

This practice test was made by AI and there may be mistakes. If a question or answer looks wrong, use the Flag a mistake button to let Brett know so he can fix it.

1 **A substance has molar mass 18 g per mole. The amount in 54 g of it, in moles, is ____**

ANSWER _____

Understand It with Pulse | Flag a mistake

2 **A substance has molar mass 100 g per mole. The mass of 2 moles of it, in grams, is ____**

ANSWER _____

Understand It with Pulse | Flag a mistake

3 **In a reaction, 3 mol of hydrogen and 1 mol of oxygen react to form water ($2H_2 + O_2 \rightarrow 2H_2O$). Which reactant runs out first?**

- A Hydrogen
- B Oxygen
- C Water
- D Neither, they run out together

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4 **The simplest whole number ratio of elements in a compound is called**

- A balanced equation
- B empirical formula
- C mole
- D molar mass

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5 **The study of amounts in chemical reactions is called**

- **A** excess reactant
- **B** percent yield
- **C** stoichiometry
- **D** empirical formula

Understand It with Pulse **Flag a mistake**

6 **6.02×10^{23} particles per mole is called**

- **A** Avogadro number
- **B** limiting reactant
- **C** excess reactant
- **D** percent yield

Understand It with Pulse **Flag a mistake**

7 **Two moles of water contain this many moles of hydrogen atoms**

- **A** 2
- **B** 1
- **C** 8
- **D** 4

Understand It with Pulse | Flag a mistake

8 **A substance has molar mass 44 g per mole. The amount in 132 g of it, in moles, is _____**

ANSWER _____

Understand It with Pulse | Flag a mistake

9 **A substance has molar mass 58.5 g per mole. The mass of 3 moles of it, in grams, is _____**

ANSWER _____

Understand It with Pulse | Flag a mistake

10 **A substance has molar mass 17 g per mole. The amount in 17 g of it, in moles, is _____**

ANSWER

Understand It with Pulse | Flag a mistake

ANSWER KEY

1.	3	**5.**	C	**9.**	175.5
2.	200	**6.**	A	**10.**	1
3.	B	**7.**	D		
4.	B	**8.**	3		

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