SASKATCHEWAN PHYSICAL SCIENCE 20 · UNIT C · 25 QUESTIONS

The Mole and Stoichiometry — Practice Test

This practice test was made by AI and there may be mistakes. If a question or answer looks wrong, use the Flag a mistake button to let Brett know so he can fix it.

1 **A substance has molar mass 18 g per mole. The amount in 36 g of it, in moles, is _____**

ANSWER _____

Understand It with Pulse | Flag a mistake

2 **A substance has molar mass 18 g per mole. The mass of 3 moles of it, in grams, is _____**

ANSWER _____

Understand It with Pulse | Flag a mistake

3 **A sample of mass 44 g contains 12 g of one element. The percent by mass of that element is _____**

ANSWER _____

Understand It with Pulse | Flag a mistake

4 **A balanced equation must have equal numbers of each atom on both sides because mass is**

A created

B destroyed

C doubled

D conserved

Understand It with Pulse | Flag a mistake

5 **A reactant left over after a reaction ends is called**

A empirical formula

B molecular formula

C excess reactant

D balanced equation

Understand It with Pulse | Flag a mistake

6 **Actual yield divided by theoretical yield, times 100 is called**

A percent yield

B molecular formula

C balanced equation

D mole

Understand It with Pulse | Flag a mistake

7 **An equation with equal numbers of each atom on both sides is called**

A molar mass

B Avogadro number

C stoichiometry

D balanced equation

Understand It with Pulse | Flag a mistake

8 **One mole of any substance contains about**

- **A** 6.02×10^{22} particles
- **B** 6.02×10^{23} particles
- **C** 6.02×10^{24} particles
- **D** 6.02×10^{20} particles

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9 **Percent yield is found from**

- **A** theoretical over actual only
- **B** actual yield over theoretical yield
- **C** mass over volume
- **D** moles over mass

Understand It with Pulse | Flag a mistake

10 **The actual number of atoms of each element in a molecule is called**

- A mole
- B molar mass
- C molecular formula
- D Avogadro number

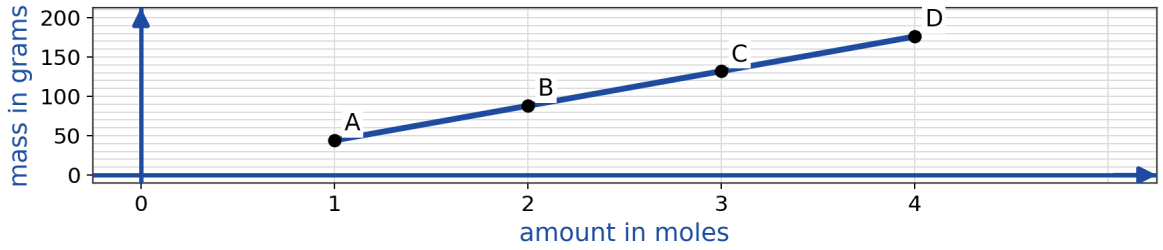
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11 **The amount containing 6.02×10^{23} particles is called**

- A mole
- B Avogadro number
- C stoichiometry
- D limiting reactant

Understand It with Pulse | Flag a mistake

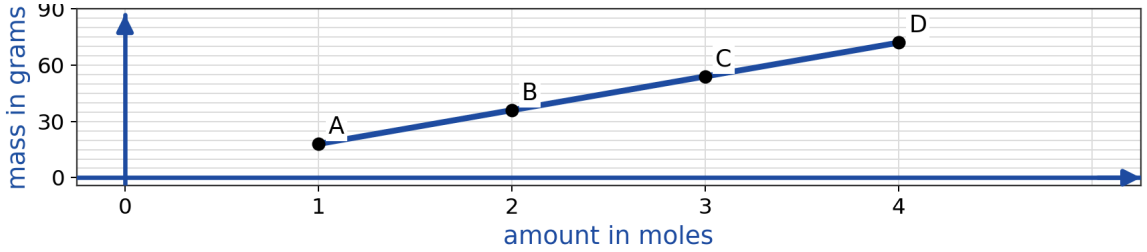
12 **The graph shows mass against amount for carbon dioxide. The mass at 2 moles, in grams, is _____**



ANSWER

Understand It with Pulse | Flag a mistake

13 **The graph shows mass against amount for water. The mass at 3 moles, in grams, is _____**



ANSWER

Understand It with Pulse | Flag a mistake

14 **The limiting reactant is the reactant that**

A is left over

B never reacts

C is a catalyst

D runs out first

Understand It with Pulse | Flag a mistake

15 **The mass of one mole of a substance is called**

A stoichiometry

B limiting reactant

C molar mass

D excess reactant

Understand It with Pulse | Flag a mistake

16 **A substance has molar mass 44 g per mole. The amount in 88 g of it, in moles, is _____**

ANSWER _____

Understand It with Pulse | Flag a mistake

17 **A substance has molar mass 58.5 g per mole. The mass of 2 moles of it, in grams, is _____**

ANSWER _____

Understand It with Pulse | Flag a mistake

18 **A sample of mass 17 g contains 14 g of one element. The percent by mass of that element is _____**

ANSWER _____

Understand It with Pulse | Flag a mistake

19 **A substance has molar mass 24 g per mole. The amount in 48 g of it, in moles, is _____**

ANSWER _____

Understand It with Pulse | Flag a mistake

20 **A substance has molar mass 17 g per mole. The mass of 4 moles of it, in grams, is ____**

ANSWER _____

Understand It with Pulse | Flag a mistake

21 **A sample of mass 18 g contains 2 g of one element. The percent by mass of that element is ____**

ANSWER _____

Understand It with Pulse | Flag a mistake

22 **A substance has molar mass 17 g per mole. The amount in 68 g of it, in moles, is ____**

ANSWER _____

Understand It with Pulse | Flag a mistake

23 **A substance has molar mass 40 g per mole. The mass of 2 moles of it, in grams, is ____**

ANSWER _____

Understand It with Pulse | Flag a mistake

24 **A substance has molar mass 100 g per mole. The amount in 100 g of it, in moles, is _____**

ANSWER _____

Understand It with Pulse | Flag a mistake

25 **A substance has molar mass 32 g per mole. The mass of 5 moles of it, in grams, is _____**

ANSWER _____

Understand It with Pulse | Flag a mistake

ANSWER KEY

1.	2	**10.**	C	**19.**	2
2.	54	**11.**	A	**20.**	68
3.	27.2727	**12.**	88	**21.**	11.1111
4.	D	**13.**	54	**22.**	4
5.	C	**14.**	D	**23.**	80
6.	A	**15.**	C	**24.**	1
7.	D	**16.**	2	**25.**	160
8.	B	**17.**	117		
9.	B	**18.**	82.3529		

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